



$$= \{ \tau^{\alpha_1} \tau^{\alpha_2} \dots \tau^{\alpha_n} \} = \{ \tau^{\alpha_1} \tau^{\alpha_2} \dots \tau^{\alpha_n} \}^J$$

$$\begin{aligned} & \text{def. } \Gamma^* = \{ \tau^{\alpha_1} \tau^{\alpha_2} \dots \tau^{\alpha_n} \}^J \\ & \text{def. } \Gamma^* = \{ \tau^{\alpha_1} \tau^{\alpha_2} \dots \tau^{\alpha_n} \}^J \\ & \text{def. } \Gamma^* = \{ \tau^{\alpha_1} \tau^{\alpha_2} \dots \tau^{\alpha_n} \}^J \\ & \text{def. } \Gamma^* = \{ \tau^{\alpha_1} \tau^{\alpha_2} \dots \tau^{\alpha_n} \}^J \end{aligned}$$

$$\text{def. } \Gamma^* = \{ \tau^{\alpha_1} \tau^{\alpha_2} \dots \tau^{\alpha_n} \}^J$$

$$f^*(\mathbf{x}, \mathbf{y}) = \frac{1}{2} \mathbf{x}^\top \mathbf{B}_1 \mathbf{B}_1^\top \mathbf{x} + \mathbf{c}^\top \mathbf{x} + \mathbf{d}^\top \mathbf{y} + \epsilon \sim \mathcal{N}(0, 100)$$

$$\frac{1}{2} \left( \frac{\partial^2}{\partial x^2} + \frac{\partial^2}{\partial y^2} \right) \psi = - \frac{1}{2} \left( \frac{\partial^2}{\partial x^2} + \frac{\partial^2}{\partial y^2} \right) \psi$$

$$V_{\text{ext}}(r) = \frac{1}{2} m \omega_r^2 r^2 + \frac{1}{2} m \omega_\theta^2 (r^2 - R_c^2)$$

$$10^{13} M_{\odot} \> 10^{13} M_{\odot}$$

$$\frac{1}{2} \left( \Gamma_{\alpha\beta}^{\mu\nu} - \Gamma_{\beta\alpha}^{\mu\nu} \right) = \frac{1}{2} \Gamma^{\mu\nu} - \frac{1}{2} \Gamma^{\nu\mu} = \frac{1}{2} \eta^{\mu\nu}$$

$$\frac{1}{2} \left( \Gamma_{\alpha\beta}^{\mu\nu} - \Gamma_{\beta\alpha}^{\mu\nu} \right) = \frac{1}{2} \Gamma^{\mu\nu} - \frac{1}{2} \Gamma^{\nu\mu} = \frac{1}{2} \eta^{\mu\nu}$$

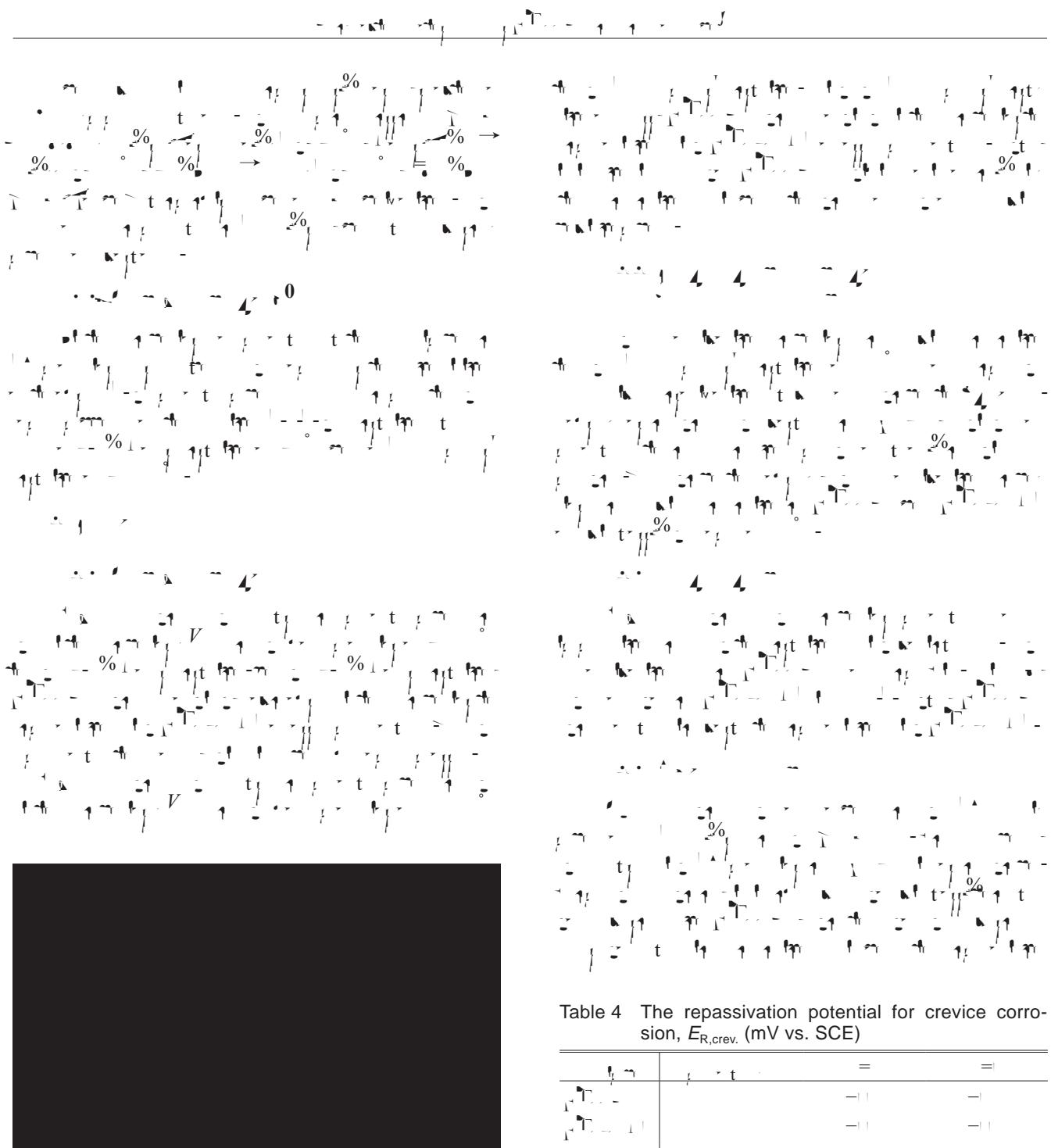


Fig.6 Pitting potential of the steels in 3.5% NaCl solutions

Fig.7 Pitting potential of the steels in 200 ppm Cl<sup>-</sup> solutions

100 200 300 400 500 600 700 800 900

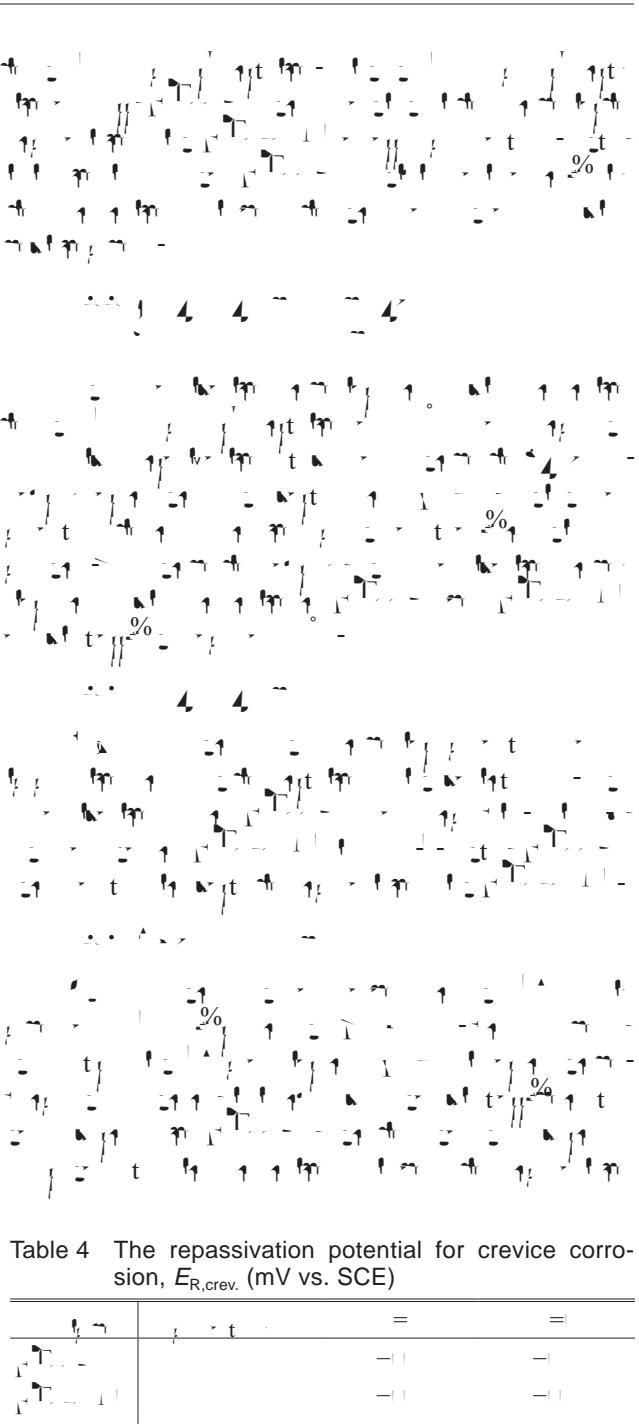


Table 4 The repassivation potential for crevice corrosion,  $E_{R,crev}$ . (mV vs. SCE)

$\text{pH}$	$E_{R,crev}$	$E_{R,crev}$
1.5	-110	-110
2.5	-110	-110
3.5	-	-

Fig.8 The rest potential after 16 hour immersion in solutions with several pH

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100% 100% 100% 100% 100% 100%  
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